



## Article

# Rational Design of an Ion-Imprinted Polymer for Aqueous Methylmercury Sorption

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**Abstract:** Methylmercury ( $\text{MeHg}^+$ ) is a mercury species that is very toxic for humans, and its monitoring and sorption from environmental samples of water are a public health concern. In this work, a combination of theory and experiment was used to rationally synthesize an ion-imprinted polymer (IIP) with the aim of the extraction of  $\text{MeHg}^+$  from samples of water. Interactions among  $\text{MeHg}^+$  and possible reaction components in the pre-polymerization stage were studied by computational simulation using density functional theory. Accordingly, 2-mercaptobenzimidazole (MBI) and 2-mercaptobenzothiazole (MBT), acrylic acid (AA) and ethanol were predicted as excellent sulfhydryl ligands, a functional monomer and porogenic solvent, respectively. Characterization studies by scanning electron microscopy (SEM) and Brunauer–Emmett–Teller (BET) revealed the obtention of porous materials with specific surface areas of  $11 \text{ m}^2 \text{ g}^{-1}$  (IIP-MBI-AA) and  $5.3 \text{ m}^2 \text{ g}^{-1}$  (IIP-MBT-AA). Under optimized conditions, the maximum adsorption capacities were  $157 \mu\text{g g}^{-1}$  (for IIP-MBI-AA) and  $457 \mu\text{g g}^{-1}$  (for IIP-MBT-AA). The IIP-MBT-AA was selected for further experiments and application, and the selectivity coefficients were  $\text{MeHg}^+/\text{Hg}^{2+}$  (0.86),  $\text{MeHg}^+/\text{Cd}^{2+}$  (260),  $\text{MeHg}^+/\text{Pb}^{2+}$  (288) and  $\text{MeHg}^+/\text{Zn}^{2+}$  (1510), highlighting the material's high affinity for  $\text{MeHg}^+$ . The IIP was successfully applied to the sorption of  $\text{MeHg}^+$  in river and tap water samples at environmentally relevant concentrations.

**Keywords:** bulk polymerization; computational modelling; environmental analysis; imprinting technology; mercury detection and removal; ion recognition; ionic imprinting polymers; sample preparation; separation science; water analysis

## 1. Introduction

Methylmercury ( $\text{MeHg}^+$ ) is one of the mercury chemical forms of highest toxicity for humans among mercury compounds [1]. Once in the human body,  $\text{MeHg}^+$  presents a high affinity to the sulfhydryl group present in proteins. This interaction can induce changes in the protein structures and a consequent loss of their functions [2]. The deleterious effects are more severe for fetuses;  $\text{MeHg}^+$

exposure can seriously impair neurological development [3]. In addition,  $\text{MeHg}^+$  is bioaccumulated and biomagnified in the marine food chain, as this specie is found in almost all marine species as a result of methylation of inorganic mercury by microorganisms present in sediments and aquatic organisms [4]. In natural water, reports estimated that  $\text{MeHg}^+$  percentage varied from 0.27 to 20% in relation to total mercury [5,6]. In this way, the consumption of contaminated water and seafood is the main source of non-occupational exposure to  $\text{MeHg}^+$ . For these reasons, the development of analytical strategies with the aim of  $\text{MeHg}^+$  determination/removal is a question of public health concern. Although several methods have been proposed for the determination and removal of  $\text{MeHg}^+$  in various samples [7–13], the synthesis and application of highly selective materials for efficient determination/removal of  $\text{MeHg}^+$  is a topic that deserves further attention.

Imprinted polymers are materials with the capacity to selectively capture molecules or metal ions and have become very popular sorbents for a wide variety of applications [14–16]. Particularly, ion-imprinted polymers (IIPs) are sorbents designed for targeting metal ions and have been recently used as solid phase extraction materials, sensors and separating membranes [17–22]. In the synthesis of IIPs, ligands containing electron-rich heteroatoms, such as nitrogen, phosphorus, sulfur and oxygen, are normally used to promote the formation of chelate with the targeted metal [23]. This three-step synthesis normally involves: (1) formation of a pre-polymerization complex with interaction of the metal ion, ligand and functional monomer (FM); (2) polymerization of the complex monomer with the addition of a cross-linker and initiator; (3) removal of the template to produce the imprinted polymer [24]. After proper removal of the metal ions from the bulk, specific cavities/binding sites able to selectively recognize the metal are generated [25].

IIPs for  $\text{MeHg}^+$  commonly require a trapping technique using a pre-polymerization mixture based on a ternary template/ligand/FM. The template is normally the analyte, using methylmercury chloride ( $\text{MeHgCl}$ ) [26–30] as the ion source, although  $\text{MeHgCl}$  is a hardly soluble salt. Ligands containing sulfhydryl groups are the more suitable ligands since mercury ions have a specific affinity to sulfur, even though ligands containing *N*-based functional groups have been found to interact with  $\text{MeHg}^+$  [24]. 1-pyrrolidinedicarbodithioic acid [26–28], phenobarbital [29–31] and methacrylic acid [26–31] have been reported in the synthesis of  $\text{MeHg}^+$  IIPs such as sulfhydryl ligand, *N*-based ligand and FM, respectively. In the present study, sulfhydryl ligands (2-mercaptobenzimidazole and 2-mercaptobenzothiazole) are tested for first time in  $\text{MeHg}^+$  IIP synthesis.

Computational simulation allows the evaluation of several components for IIP synthesis and their interaction; it is assumed that the formed complexes preserve their chemical structures after polymerization, thus, only the pre-polymerization system is modeled [32,33]. Therefore, the consideration of individual interactions among template, ligand, FM, cross-linker and solvent might significantly reduce the time and resources to choose the components for IIP synthesis. Density functional theory (DFT) methods are typically used to calculate the “binding energy” for polymer components selection [34,35].

In this work, theoretical calculations were used to predict relevant experimental conditions in the synthesis of an IIP for aqueous  $\text{MeHg}^+$  sorption. DFT was used for geometry optimization and binding energy calculations in systems containing  $\text{MeHg}^+$ , sulfhydryl ligands, FM and solvent. For the validation of the theoretical results, IIPs and non-imprinted polymers (NIPs) were synthesized by bulk polymerization and characterized by Fourier transform infrared spectroscopy (FTIR), scanning electron microscopy (SEM), thermogravimetric analysis (TGA) and Brunauer–Emmett–Teller (BET) measurements. The sorption and selectivity study were performed to assess the efficiency of the proposed  $\text{MeHg}^+$ -IIP and for comparison with the corresponding NIPs. Finally, the  $\text{MeHg}^+$ -IIP was successfully applied to the quantitative sorption of  $\text{MeHg}^+$  in samples in river and tap water at environmentally relevant concentrations.

## 2. Materials and Methods

### 2.1. Reagents and Samples

MeHgCl, tiourea, 2-mercaptobenzimidazole (MBI), 2-mercaptobenzothiazole (MBT), acrylic acid (AA), ethylene glycol dimethacrylate (EGDMA) and azodiisobutyronitrile (AIBN) were purchased from Sigma-Aldrich (St. Louis, MO, USA); ethanol (EtOH) absolute (99.9%) was obtained from JT Baker (Radnor, PA, USA); MeHgCl standard solution 1000 mg L<sup>-1</sup> from Alfa Aesar (Tewksbury, MA, USA); and hydrochloric acid (HCl), sodium acetate (NaCH<sub>3</sub>COO), acetic acid (CH<sub>3</sub>COOH), sodium hydrogen phosphate (Na<sub>2</sub>HPO<sub>4</sub>), sodium dihydrogen phosphate (NaH<sub>2</sub>PO<sub>4</sub>), sodium hydroxide (NaOH), and sodium tetraborate (Na<sub>2</sub>B<sub>4</sub>O<sub>7</sub>) were purchased from Merck (Darmstadt, Germany). For the selectivity study, solutions were prepared from salt standards of lead nitrate, cadmium nitrate tetrahydrate, zinc sulfate monohydrate, and mercury dichloride (all acquired from Sigma-Aldrich). All the reagents used in this work were of analytical grade, and the solutions were prepared in ultrapure water (resistivity of 18.2 MΩ cm).

The samples were collected from the Batalha river (21.72 S, 49.20 W), and directly from the tap.

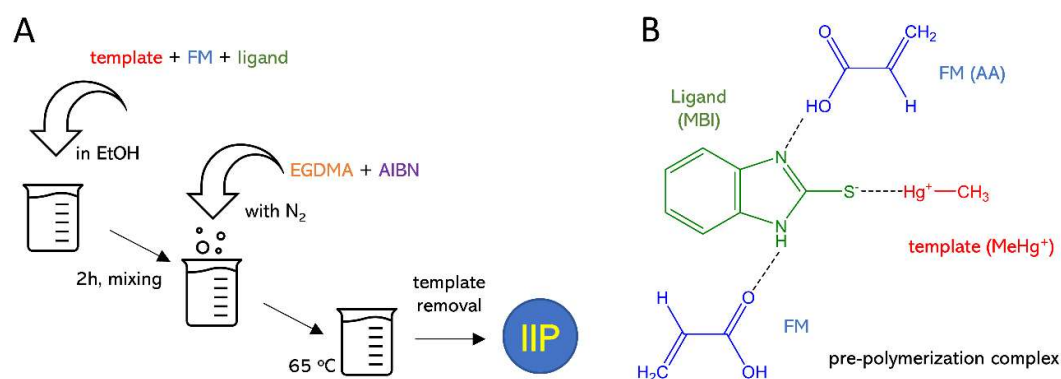
WARNING: MeHgCl appears as white microcrystals and is very toxic by inhalation, in contact with skin and if swallowed. Protection measures such as glasses, full face-shields, gloves and particulate respirators are required to manipulate it. Its disposal should also be carried out carefully.

### 2.2. Density Functional Method

DFT is a computational method applied to electronic systems to obtain the energy and the electronic distribution of the fundamental state, which is a simpler magnitude than the wave function and therefore simplifies the study of complex systems [33,34]. Accordingly, the binding energies were calculated using Gibbs free energies ( $\Delta G$ ), which is a compressive parameter (involving enthalpy, entropy and temperature) and a determinant thermodynamic criterion for a spontaneous process [32]. The computational simulations were performed using a Gaussian 09 package [36], M06-2X level of theory [37] and basis set 6-31G(d) [38,39] for the carbon, sulfur, nitrogen, oxygen and hydrogen atoms. In the case of MeHg<sup>+</sup>, a pseudopotential (LANL2DZ) [40] was employed to describe the core orbitals of the mercury atom.

### 2.3. Ion-Imprinted Polymer (IIP) Synthesis

MeHg<sup>+</sup>-IIPs were synthesized via bulk polymerization using MeHgCl (0.5 mmol) as the template; MBI (0.5 mmol) or MBT (0.5 mmol) as ligands for comparison; AA (1.0 mmol) as FM; EGDMA (2 mmol) as a cross-linker, AIBN (0.5 mmol) as radical initiator; EtOH as solvent (20 mL). First, the template and ligand were dissolved in EtOH, the FM was added, and the mixture was kept under vigorous stirring for 2 h. Subsequently, the cross-linker and the radical initiator were added under nitrogen atmosphere. Then, the glass tube was sealed under an inert environment, and the temperature was set to 65 °C to initiate the polymerization. A schematic illustration of the IIP synthesis is shown in Figure 1.



**Figure 1.** Schematic illustration of the synthesis of ion-imprinted polymers (IIPs) (A); pre-polymerization complex with the 2-mercaptobenzimidazole (MBI) ligand (B).

The template ion  $\text{MeHg}^+$  in the IIPs was removed with a mixture of thiourea 2.5% *w/v* and  $\text{HCl}$  1 mol  $\text{L}^{-1}$  solution until mercury could no longer be detected in the solution by thermal decomposition amalgamation atomic absorption spectrometry (TDA–AAS), following a previously proposed method [10]. The polymer was dried at 60 °C and stored at room temperature. Two non-imprinted polymers (NIPs) were synthesized following the same procedure, but no template was added in one of them, and neither template nor ligand was added in the other. Thus, the synthesized materials were IIP–MBI–AA and IIP–MBT–AA ( $\text{CH}_3\text{Hg}^+$  imprinted polymers), NIP–MBI–AA and NIP–MBT–AA (non-imprinted polymers with ligand) and NIP–AA (non-imprinted polymers without ligand).

#### 2.4. Polymer Characterization

Characterizations of the  $\text{MeHg}^+$ –IIPs and NIPs were carried out by attenuated total reflection (ATR)–FTIR in a Bruker spectrometer (Billerica, MA, USA) Alpha II in the range of 4000–400  $\text{cm}^{-1}$  to identify functional groups in each material. Morphology was examined by scanning electron microscopy (SEM) of the JEOL brand model JSM7500F (Tokyo, Japan). The samples were dried at 60 °C, placed on a carbon wafer and then subjected to a sputtering using a spray coating equipment of the Bal Tec brand model SCD 050. The nitrogen adsorption/desorption analysis was used to evaluate the surface area by the BET method (Brunauer–Emmett–Teller), whereas the total volume and pore diameter was evaluated by the BJH method (Barret–Joyner–Halenda) using a Micromeritics Gemini VII-2390 (Norcross, GA, USA). The samples were pretreated in a vacuum by degassing at 100 °C for 2 h. The thermogravimetric analysis (TGA) and differential thermal analysis (DTA) were applied to evaluate the thermal stability of the materials. Instrument Perkin Elmer STA 6000 (Waltham, MA, USA) was used to obtain the thermograms. The thermograms were obtained using 5 mg of mass of the sample and a heating ramp of 10 °C  $\text{min}^{-1}$  in the range of from 35 to 800 °C in nitrogen atmosphere with a flow rate of 20  $\text{mL min}^{-1}$  using a Perkin Elmer STA 6000 instrument.

#### 2.5. Sorption Studies

Sorption experiments of  $\text{MeHg}^+$  from aqueous solutions were carried out by TDA–AAS, using a direct mercury analyzer (DMA) for the determination of total mercury [10]. Sorption isotherms were studied at different  $\text{MeHg}^+$  concentrations (25, 50, 100, 150, 200, 300, 400, 600 and 800  $\text{mg L}^{-1}$ ). Polymer concentration was kept constant at 3 mg of polymeric material in 2 mL of  $\text{MeHg}^+$  solution and shaken for 2 h. Then, the polymers were separated by centrifugation for 10 min at 450 RCF. After that, the concentration of  $\text{MeHg}^+$  was quantified in the supernatant using TDA–AAS. The effect of pH on  $\text{MeHg}^+$  sorption was also studied with the following buffer solutions:  $\text{HCl}$  for pH 2.0;  $\text{NaCH}_3\text{COO}/\text{CH}_3\text{COOH}$  for pH 4.0–6.0;  $\text{Na}_2\text{HPO}_4/\text{NaH}_2\text{PO}_4$  for pH 7.0–8.0;  $\text{Na}_2\text{B}_4\text{O}_7/\text{NaOH}$  for pH 10.

The sorption kinetic was studied in the range of 5–300 min. Afterwards, the samples were centrifuged for 10 min at 450 RCF, and the residual solution was analyzed by TDA–AAS. The other parameters were kept constant (100  $\mu\text{g L}^{-1}$   $\text{MeHg}^+$  solution, pH 8 with  $\text{Na}_2\text{HPO}_4/\text{NaH}_2\text{PO}_4$  buffer, polymer concentration: 3 mg of polymeric material in 2 mL  $\text{MeHg}^+$  solutions, and 2 h of stirring time).

#### 2.6. Selectivity Studies and Water Analysis

The selectivity of the synthesized polymer for  $\text{MeHg}^+$  was assessed in binary mixtures of  $\text{MeHg}^+/\text{Hg}^{2+}$ ,  $\text{MeHg}^+/\text{Cd}^{2+}$ ,  $\text{MeHg}^+/\text{Pb}^{2+}$  and  $\text{MeHg}^+/\text{Zn}^{2+}$ . A polymer mass of 3 mg and 2 mL standard  $\text{MeHg}^+$  solutions at pH 8 were used for these experiments under shaking for 2 h. The mixture was centrifuged for 10 min at 450 RCF, and the concentration of metal ions was quantified using inductively coupled plasma with optical emission spectrometry (ICP–OES; Optima 3300 DV from Perkin Elmer). The same procedure was used to analyze the water samples. To evaluate the application of the materials, the two water samples were analyzed with the same procedure. These were filtered using Whatman #42 paper and tested by spiking and recovery experiments. The  $\text{MeHg}^+$  in the IIP was removed with a mixture of thiourea, 2.5% *w/v*, and 1 mol  $\text{L}^{-1}$  of  $\text{HCl}$  following a previously proposed methodology [27], and it was then quantified by TDA–AAS [10].

### 3. Results and Discussion

#### 3.1. Theoretical Selection of the Functional Monomer and Solvent

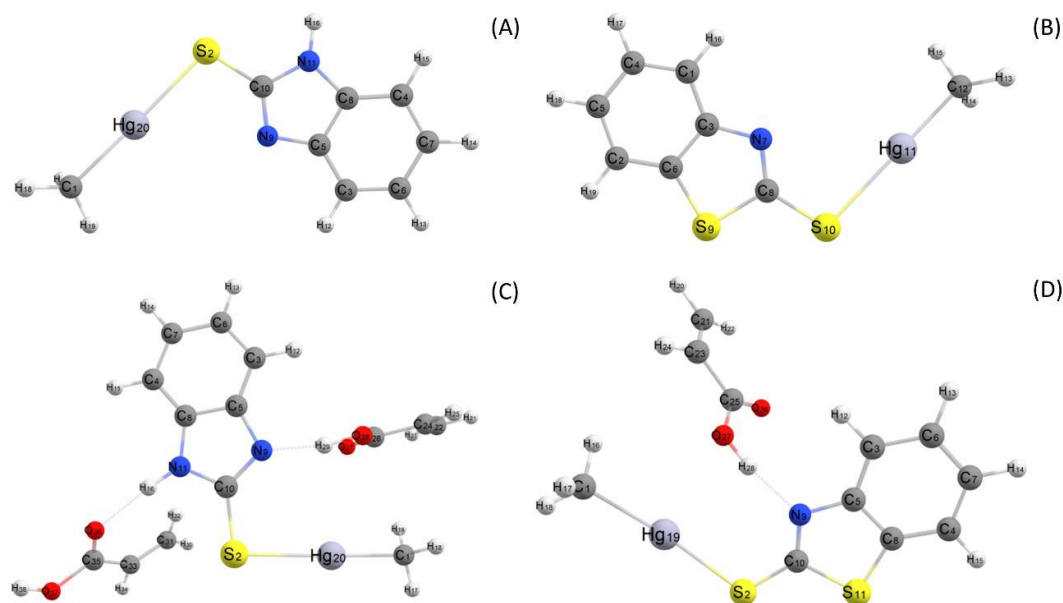
A methodology based on the geometry optimization was employed to obtain the lowest energy conformation followed by the calculation of the binding energy ( $\Delta E_1$ ) of each metal ion–ligand complex. This process was also carried out for the metal ion–ligand–FM complex ( $\Delta E_2$ ). The interaction energies of complexes were based on the following equations (where  $n$  is number of FM for metal ion–ligand–FM complex calculations):

$$\Delta E_1 = E(\text{metal ion} - \text{ligand}) - E(\text{metal ion}) - E(\text{ligand}), \quad (1)$$

$$\Delta E_2 = E(\text{metal ion} - \text{ligand} - \text{FM}) - E(\text{metal ion} - \text{ligand}) - nE(\text{FM}), \quad (2)$$

The assessment of these energies in different solvents (water, EtOH, dimethyl sulfoxide (DMSO), dimethylformamide (DMF) and acetonitrile) was also carried out aiming to predict the most stable complex.

The geometry optimization was performed for  $\text{MeHg}^+$  complexes with the sulfhydryl ligands 2-mercaptobenzimidazole (MBI) and 2-mercaptobenzothiazole (MBT). As the linear coordination is the most likely geometry for  $\text{MeHg}^+$  complexes [41], the formation of linear complexes of  $\text{MeHg}^+$ –MBI and  $\text{MeHg}^+$ –MBT was considered in the geometry optimization in EtOH medium (see Figure 2). For theoretical calculations, the Hg–S distances were calculated to be 2.717 Å and 2.698 Å for  $\text{MeHg}^+$ –MBT and  $\text{MeHg}^+$ –MBI, respectively, which were similar to the experimental values previously reported [42]. Some additional theoretical and experimental bond lengths for these complexes can be found in Table S1.



**Figure 2.** Geometry optimization of methylmercury ( $\text{MeHg}^+$ )–MBI (A) and  $\text{MeHg}^+$ –2-mercaptobenzothiazole (MBT) (B) in ethanol (EtOH), and optimized geometries of the pre-polymerization step of  $\text{MeHg}^+$ –MBI (C) and  $\text{MeHg}^+$ –MBT (D) complexes with acrylic acid (AA) as the functional monomer (FM).

The effect of the solvent on the optimized geometries was also considered, and the binding energies for the  $\text{MeHg}^+$  complexes are presented in Table 1. The results display negative value of  $\Delta E$ , which indicates spontaneous complex formation and an effective interaction between the Hg atom and sulfhydryl groups. The most stable (lowest) values obtained were  $-15.4 \text{ kcal mol}^{-1}$  for  $\text{MeHg}^+$ –MBI and  $-12 \text{ kcal mol}^{-1}$  for  $\text{MeHg}^+$ –MBT in EtOH, and it was therefore selected as the porogenic solvent.



**Table 1.** Solvent effect on binding energy in MeHg<sup>+</sup> complexes. The dielectric constant values for water, EtOH, dimethyl sulfoxide (DMSO), dimethylformamide (DMF) and acetonitrile were 78.4, 24, 47, 36.7 and 36.5, respectively.

Complex	Binding Energy/kcal mol <sup>−1</sup>					
	Water	EtOH	DMSO	DMF	Acetonitrile	Vacuum
MeHg <sup>+</sup> –MBI	−10.5	−15.4	−12.2	−13.1	−13.3	−168.5
MeHg <sup>+</sup> –MBT	−7.8	−12.0	−9.2	−10.0	−10.2	−161.5

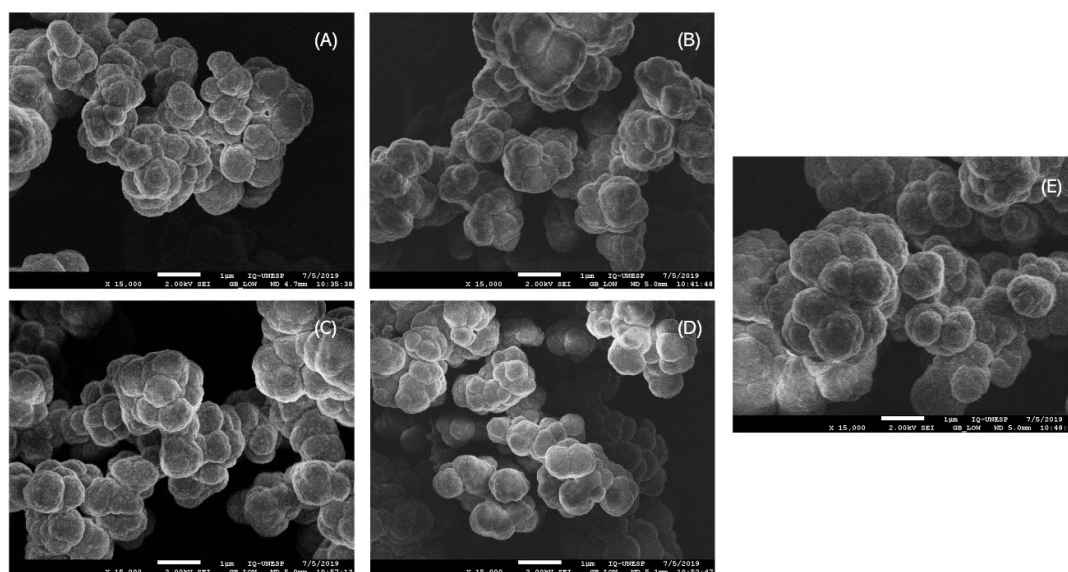
Regarding the FM, it should possess higher binding energy toward the ligand than the binding energy with the template to allow for the template/analyte to be adsorbed and desorbed from the IIP bulk. Figure 2 and Table S2 show the chemical structures and the  $\Delta E$  of the pre-polymerization stage in the EtOH solvent. These results show that AA provides stronger IIPs than basic and neutral monomers. MeHg<sup>+</sup> interacts with MBI (−13.3 kcal mol<sup>−1</sup>) and MBT (−11 kcal mol<sup>−1</sup>) by coordinate bonds (1:1), and AA interacts with MBI (−20 kcal mol<sup>−1</sup>) and MBT (−12.2 kcal mol<sup>−1</sup>) by hydrogen bonds. Thus, after the theoretical selection of two potentially efficient ligands (MBI and MBT), the FM (AA) and solvent (EtOH), these results were experimentally validated by synthesizing and comparing of the performances of the IIP–MBI–AA and IIP–MBT–AA materials.

### 3.2. Characterization of the Synthesized Polymers

Figure S1 shows the comparison of the FTIR spectra of all the synthesized polymers, which presented similar spectral profiles. IIPs and NIPs showed characteristic peaks of C=O stretch (1724 cm<sup>−1</sup>), C–O stretch (1146 cm<sup>−1</sup>), C–O bend (1253 cm<sup>−1</sup>), C–H stretch (2980 cm<sup>−1</sup>) [43], CH<sub>3</sub> and CH<sub>2</sub> bends (1392 and 1452 cm<sup>−1</sup>, respectively). As expected, no band is present in the region of 1645–1630 cm<sup>−1</sup> of C=C groups in polymer materials, suggesting the complete polymerization of EGDMA. The sulfhydryl ligands MBI and MBT were also characterized by FTIR; however, the characteristic peaks associated with MBI (N–C=S stretch (1174 cm<sup>−1</sup>) and S–H stretch (2877 cm<sup>−1</sup>)) and with MBT (NS–H stretch (2890 cm<sup>−1</sup>), N–C=S stretch (1082 cm<sup>−1</sup>) and C–S stretch (748–667 cm<sup>−1</sup>)) were not observed [44,45]. It should be noted that, the FTIR spectroscopy analysis was not able to provide vibrational information about the ligands in the polymer network because of the small amount of MBI or MBT used in the reaction compared to the cross-linker (EGDMA).

The surface morphology of the synthesized materials was studied by SEM analysis, and the results are shown in Figure 3. In all the cases, rounded particles were obtained, and a slightly higher roughness for the IIP–MBT–AA and NIP–MBT–AA materials was observed. As there were no clear difference among the morphology of the synthesized polymers by SEM analysis, the quantitative measurement of the effective surface area was necessary.

Nitrogen adsorption/desorption analysis was carried out to estimate the pore surface area by BET, and the pore size and total pore volume by Barret–Joyner–Halenda (BJH) analysis. The results are presented in the Table 2, and it can be observed that IIP–MBI–AA presented the largest surface area (11 m<sup>2</sup> g<sup>−1</sup>) and the smallest pore size (9.5 nm). This suggests the presence of more imprinted sites for MeHg<sup>+</sup> and a high adsorption capacity for the IIP–MBI–AA polymer. The IIP–MBT–AA and NIP–MBT–AA materials displayed similar surface areas (5.3 and 5.5 m<sup>2</sup> g<sup>−1</sup>, respectively) and pore diameter (11 and 12 nm, respectively), which indicates a small difference on the sorption capacity. Based on this data, IIP–MBI–AA might be considered as a better sorbent than IIP–MBT–AA since the MBI ligand interacts with two hydrogen bonds in the polymer network, creating more binding sites to ensure migration of MeHg<sup>+</sup>. On the other hand, the MBT ligand only interacts with one hydrogen bond, indicating lower porosity towards binding sites. However, quantitative sorption studies are necessary to corroborate these hypotheses.



**Figure 3.** SEM images of (A) IIP-MBI-AA, (B) non-imprinted polymer (NIP)-MBI-AA, (C) IIP-MBT-AA, (D) NIP-MBT-AA, and (E) NIP-AA; the scale bar is 1  $\mu\text{m}$  in all figures.

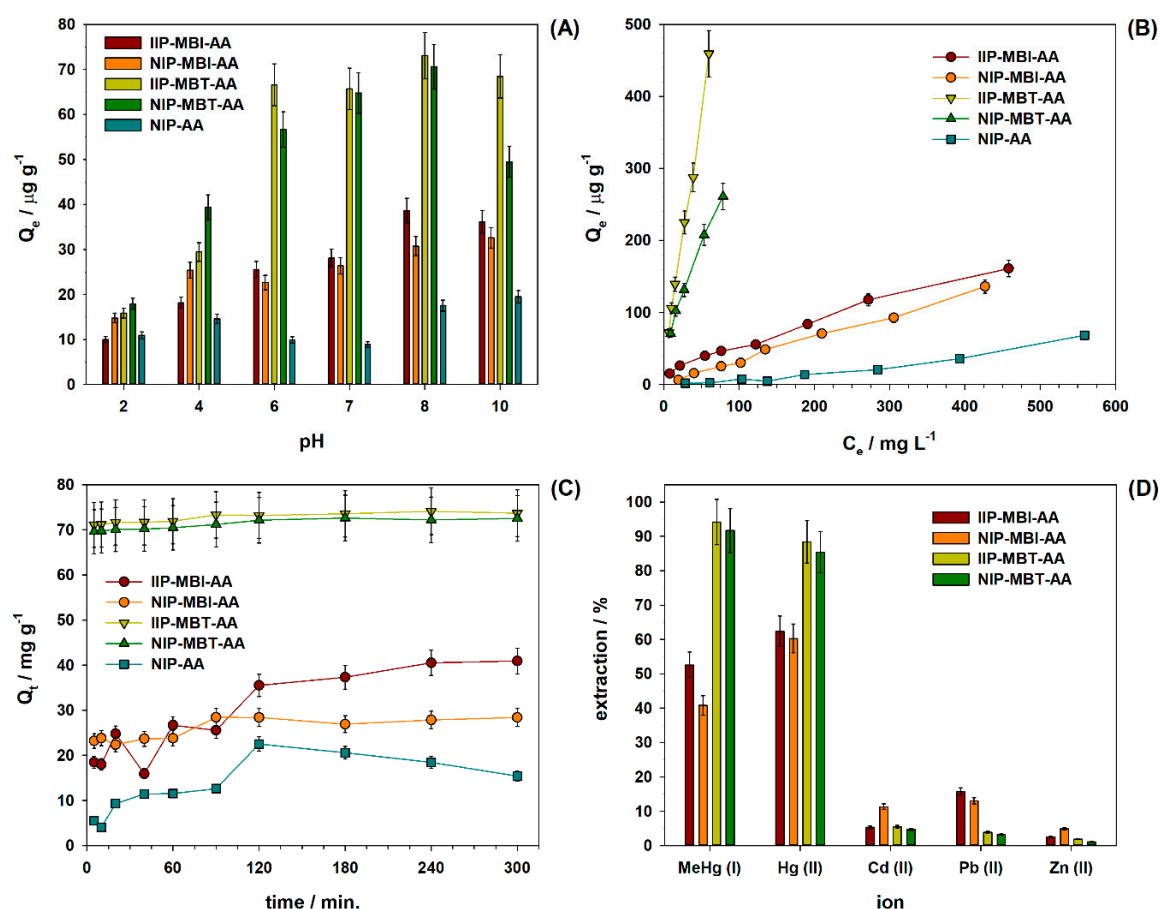
**Table 2.** Calculated Brunauer–Emmett–Teller (BET) surface area, total pore volume and average pore size.

Material	BET Surface Area/ $\text{m}^2 \text{g}^{-1}$	Pore Size/nm
IIP-MBI-AA	11	9.5
NIP-MBI-AA	6.8	21
IIP-MBT-AA	5.3	17
NIP-MBT-AA	5.5	11
NIP-AA	3.0	12

Thermal stability of the polymers was evaluated by TGA and DTA techniques, and the results can be found in Figure S2. All polymers showed similar degradation pattern. The first endothermic event is observed due to the loss of intrinsically bound water at a temperature below 100  $^{\circ}\text{C}$  in the TGA/DTA curve. This peak is followed by another endothermic peak which is responsible for the main decomposition of the polymers (282–333  $^{\circ}\text{C}$ ). It could be that this is attributed to the short chain degradations as well as decarboxylation process, and was continued until 400  $^{\circ}\text{C}$  [46–48]. Nearly 100% of total mass of the polymers was decomposed in the range of 282–400  $^{\circ}\text{C}$ . This means that all polymers possess good thermal stability, tolerating temperatures until 282  $^{\circ}\text{C}$  without significant degradation.

### 3.3. Adsorption Studies

The effect of pH on  $\text{MeHg}^+$  sorption towards IIPs and NIPs was studied in the pH range from 2 to 10, and the results are presented in Figure 4A. In all materials, the sorption capacity increased rapidly and reached the maximum capacity at pH 8. For IIP-MBI-AA, the maximum value was 39  $\mu\text{g g}^{-1}$ , whereas for NIP-MBI-AA and NIP-AA, the maximum were 30  $\mu\text{g g}^{-1}$  and 17  $\mu\text{g g}^{-1}$ , respectively. On the other hand, IIP-MBT-AA (73  $\mu\text{g g}^{-1}$ ) and NIP-MBT-AA (70  $\mu\text{g g}^{-1}$ ) display similar adsorption capacity. These results indicate that despite the fact that the IIP-MBI-AA polymer displayed the larger surface area, the IIP-MBT-AA polymer exhibited a highest adsorption capacity. This was probably due to the larger pore size of IIP-MBT-AA which facilitates the analyte sorption. The pH has a strong influence on  $\text{MeHg}^+$  speciation; for example, a pH around 6–8 dictates that most of the  $\text{MeHg}^+$  would have a hydroxide group associated with it, while in low pH 2–4, ionic speciation for  $\text{MeHg}^+$  is present. pH 8 was selected to perform all the following experiments because the sorption was more efficient at this pH value. The same trend was observed in another IIP for  $\text{MeHg}^+$  where the binding capacities increased with high values of pH [30].



**Figure 4.** (A) Effect of pH on sorption of  $\text{MeHg}^+$  on IIP-MBI-AA, NIP-MBI-AA, IIP-MBT-AA, NIP-MBT-AA, and NIP-AA; 3 mg of polymeric material,  $[\text{MeHg}^+] = 100 \mu\text{g L}^{-1}$ , volume = 2.0 mL, shaking time = 2 h, room temperature. (B) The effect of  $\text{MeHg}^+$  initial concentration on the adsorption of  $\text{MeHg}^+$  on IIP-MBI-AA, NIP-MBI-AA, IIP-MBT-AA, NIP-MBT-AA, and NIP-AA; 3 mg of polymeric material,  $[\text{MeHg}^+] = 25\text{--}800 \mu\text{g L}^{-1}$ , pH = 8, volume = 2.0 mL, shaking time = 2 h, room temperature. (C) Kinetics of  $\text{MeHg}^+$  adsorption on IIP-MBI-AA, NIP-MBI-AA, IIP-MBT-AA, NIP-MBT-AA, and NIP-AA; 3 mg of polymeric material,  $[\text{MeHg}^+] = 100 \mu\text{g L}^{-1}$ , pH = 8, shaking time = 5–300 min, volume = 2.0 mL, room temperature. (D) The selectivity profiles of the IIPs and NIPs for binary mixtures containing  $\text{MeHg}^+$  and potential interferences. In all experiments  $n = 3$ .

Equilibrium data, also known as the adsorption isotherm, are the basic requirement for designing an adsorption system. To evaluate the adsorption capacity, we equilibrated IIPs or NIPs (3 mg) with  $\text{MeHg}^+$  solutions (2 mL) in the concentration range of  $25\text{--}600 \mu\text{g L}^{-1}$  at pH 8 (Figure 4B). The adsorption value increased with the increase in the concentration of  $\text{MeHg}^+$  in all materials, and it revealed that IIP-MBI-AA ( $161 \mu\text{g g}^{-1}$ ) has higher adsorption capacity than that of non-imprinted polymers NIP-MBI ( $135 \mu\text{g g}^{-1}$ ) and NIP-AA ( $68 \mu\text{g g}^{-1}$ ). The significant difference in adsorption capacities indicates the role of ion imprinting in the adsorbent IIP-MBI. In the same way, IIP-MBT-AA ( $457 \mu\text{g g}^{-1}$ ) and NIP-MBT-AA ( $416 \mu\text{g g}^{-1}$ ) display differences in adsorption capacity, as expected.

The data obtained were converted to Langmuir and Freundlich isotherms, and the related constants are given in Table 3. The Langmuir model is represented in a linear equation form as [49]:

$$\frac{C_e}{Q_e} = \frac{C_e}{Q_m} + \frac{1}{bQ_m}, \quad (3)$$

where  $C_e$  is equilibrium concentration of  $\text{MeHg}^+$  ( $\mu\text{g L}^{-1}$ ),  $Q_e$  is the adsorption capacity of  $\text{MeHg}^+$  adsorbed at equilibrium ( $\mu\text{g g}^{-1}$ ) and  $Q_m$  and  $b$  are Langmuir constants, which are related to the



adsorption capacity and energy of adsorption, respectively. The plot of  $C_e/Q_e$  against  $C_e$  was used to support the Langmuir isotherm.

**Table 3.** Langmuir and Freundlich isotherm constants.

Material	Langmuir			Freundlich		
	$Q_m/\mu\text{g g}^{-1}$	$b/\text{L } \mu\text{g}^{-1}$	$R^2$	$K_f/\mu\text{g g}^{-1}$	$n^{-1}/\text{L } \mu\text{g}^{-1}$	$R^2$
IIP-MBI-AA	217.4	0.004	0.76	4.281	0.572	0.98
IIP-MBT-AA	1250	0.009	0.84	13.02	0.866	0.99

The Freundlich isotherm is an exponential equation that assumes the sorbate concentration increases together with the adsorbent surface [50]. The linear equation used to study the Freundlich isotherm was

$$\log Q_e = \log K_f + \frac{1}{n} \log C_e, \quad (4)$$

where  $K_f$  is the intercept showing the sorption capacity of the adsorbents, and  $n^{-1}$  is the slope showing the variation of the sorption with the concentration. The results showed that the Freundlich model fitted ( $R^2 = 0.98\text{--}0.99$ ) better than the Langmuir model ( $R^2 = 0.76\text{--}0.84$ ) with the experimental behavior of the sorption process. This suggests the interactions and adsorption occur on multiple layers of the sorbents, which is an expected behavior for non-covalently-imprinted polymers.

To study the kinetic of  $\text{MeHg}^+$  sorption onto IIPs and NIPs, batch experiments were carried out. Figure 4C shows that  $\text{MeHg}^+$  adsorptions reach the maximum capacity after around 120 min for IIP-MBI-AA, NIP-MBI-AA and NIP-AA, and 95 min for IIP-MBT-AA and NIP-MBT-AA. The kinetic data were fitted using the pseudo-first-order and pseudo-second-order kinetic models, described by Equations (5) and (6), respectively [35,51].

$$\log(Q_e - Q_t) = \log Q_e - \frac{K_1}{2.303} t, \quad (5)$$

$$\frac{t}{Q_t} = \frac{1}{K_2 Q_e^2} + \frac{1}{Q_e} t, \quad (6)$$

where  $Q_e$  ( $\mu\text{g g}^{-1}$ ) and  $Q_t$  ( $\mu\text{g g}^{-1}$ ) are the amounts of adsorbate adsorbed at equilibrium and at time  $t$ , respectively.  $K_1$  ( $\text{min}^{-1}$ ) and  $K_2$  ( $\text{g } \mu\text{g}^{-1} \text{min}^{-1}$ ) are the pseudo-first-order and pseudo-second-order adsorption rate constants, respectively. The pseudo-first-order and pseudo-second-order kinetic models were used to model the experimental data. From Table 4, we can conclude that the pseudo-second-order kinetic model ( $R^2 \leq 1$ ) provided a good correlation for the adsorption of  $\text{MeHg}^+$  on IIPs and NIPs compared to the pseudo-first-order ( $R^2 \ll 1$ ) kinetic model. Besides, the theoretical values ( $Q_{the}$ ) calculated through the pseudo-second-order model for all polymers were close to the experimental values ( $Q_{exp}$ ).

In order to investigate and to compare the selectivity of IIP-MBI-AA and IIP-MBT-AA for  $\text{MeHg}^+$ , some metal ions ( $\text{Hg}^{2+}$ ,  $\text{Cd}^{2+}$ ,  $\text{Pb}^{2+}$ , and  $\text{Zn}^{2+}$ ) were chosen as competitive metals to prepare binary mixtures. The ion  $\text{Hg}^{2+}$  was selected as a potential interferent because it often coexists with  $\text{MeHg}^+$  in wastewater and biological samples. The ion  $\text{Cd}^{2+}$  was chosen because of its high affinity for thiol groups. The  $\text{Pb}^{2+}$  and  $\text{Zn}^{2+}$  metal ions were chosen as competitive ions because of their presence in environmental samples. The adsorption capacity ( $Q$  in  $\mu\text{g g}^{-1}$ ) and the extraction ( $E$ ) were calculated using the following equations [52]:

$$Q = \frac{(C_0 - C_e)}{m} V, \quad (7)$$

$$E = \frac{(C_0 - C_e)}{C_0}, \quad (8)$$

where  $m$  is the mass of polymer (mg),  $V$  is the volume of the solution (mL), and  $C_0$  and  $C_e$  are the initial and equilibrium concentrations of  $\text{MeHg}^+$  ( $\mu\text{g L}^{-1}$ ), respectively. Figure 4D shows the extraction percentage (obtained by Equation (4)) towards  $\text{MeHg}^+$  and other metal ions. It is evidenced that IIPs and NIPs have selectivity towards mercurial compounds due to sulfhydryl groups in their polymer structures. For  $\text{MeHg}^+$  sorption, the extraction values were 53% and 41% for IIP-MBI-AA and NIP-MBI-AA, respectively. However, higher extraction values were obtained (62% and 60% for IIP-MBI-AA and NIP-MBI-AA) for  $\text{Hg}^{2+}$  sorption. This suggests that these materials have a higher affinity towards  $\text{Hg}^{2+}$  ions than  $\text{MeHg}^+$  ions. Differently,  $\text{MeHg}^+$  extraction values were 94% and 92% for IIP-MBT-AA and NIP-MBT-AA, and  $\text{Hg}^{2+}$  extraction values were 88% and 85% for IIP-MBT-AA and NIP-MBT-AA, respectively.

**Table 4.** Pseudo-first-order and pseudo-second-order adsorption rate constants.

Pseudo-First-Order				
Material	$Q_{exp}/\mu\text{g g}^{-1}$	$Q_{the}/\mu\text{g g}^{-1}$	$K_1/\text{min}^{-1} \times 10^{-3}$	$R^2$
IIP-MBI-AA	40.93	33.23	15.43	0.886
NIP-MBI-AA	28.42	8.72	15.89	0.816
NIP-AA	22.52	14.54	4.606	0.582
IIP-MBT-AA	74.13	2.75	7.369	0.853
NIP-MBT-AA	72.62	3.40	11.28	0.916
Pseudo-Second-Order				
Material	$Q_{exp}/\mu\text{g g}^{-1}$	$Q_{the}/\mu\text{g g}^{-1}$	$K_2/\text{mg g}^{-1} \text{min}^{-1} \times 10^{-3}$	$R^2$
IIP-MBI-AA	40.93	44.44	0.678	0.958
NIP-MBI-AA	28.42	28.57	7.239	0.998
NIP-AA	22.52	18.48	2.766	0.941
IIP-MBT-AA	74.13	74.07	15.85	0.999
NIP-MBT-AA	72.62	72.46	14.21	0.999

Table 3 expresses the distribution coefficient ( $K_d$ ), the selectivity coefficient ( $k$ ) and the impression factor ( $I$ ) values, which could be obtained from the following equations:

$$K_d = \frac{(C_0 - C_e)}{m C_e} V, \quad (9)$$

$$k = \frac{K_d(\text{MeHg}^+)}{K_d(\text{ion})}, \quad (10)$$

$$I = \frac{k(\text{IIP})}{k(\text{NIP})}, \quad (11)$$

The  $K_d$  values of IIP-MBI-AA and IIP-MBT-AA for  $\text{MeHg}^+$  were higher compared to those of the respective NIPs. The  $k$  for  $\text{MeHg}^+/\text{Hg}^{2+}$ ,  $\text{MeHg}^+/\text{Cd}^{2+}$ ,  $\text{MeHg}^+/\text{Pb}^{2+}$  and  $\text{MeHg}^+/\text{Zn}^{2+}$  in IIP-MBI-AA were 0.89, 20, 5.8 and 35, respectively. On the other hand, the IIP-MBT-AA  $k$  values were 0.86, 261, 288 and 1510. These results showed that the imprinted polymers retain both species,  $\text{Hg}^{2+}$  and  $\text{MeHg}^+$ , with higher binding capacities. IIP-MBT-AA proved higher affinity than IIP-MBI-AA, and it could be because MBT has two sulfhydryl groups, while MBI only has one sulfhydryl group for complexing  $\text{MeHg}^+$  or  $\text{Hg}^{2+}$ . Taking into consideration the superior extraction efficiency of IIP-MBT-AA, and despite the slightly lower selectivity coefficient, this material was selected for  $\text{MeHg}^+$  sorption in water samples.

### 3.4. Application in Water Samples

The applicability of the IIP-MBT-AA for preconcentration of trace levels of  $\text{MeHg}^+$  was used as a solid phase adsorbent material. It was tested by spiking and recovery experiments on river water and tap water. The water samples spiked with  $\text{MeHg}^+$  at three different environmentally relevant concentrations (25, 50 and  $100 \mu\text{g L}^{-1}$ ) were treated with the imprinted polymer. The recoveries obtained were between 84% and 95% (Table 5), and the relative standard deviation varied in the range of from 1 to 10%. Recoveries were calculated by the found  $\text{MeHg}^+$  divided per added  $\text{MeHg}^+$ . These results indicate the suitability of the present  $\text{MeHg}^+$  imprinted polymer (IIP-MBT-AA) for the quantitative extraction of  $\text{MeHg}^+$  from natural water samples taking into account the complexity of matrix on the river water sample with recovery percentages slightly lower than tap water.

**Table 5.** Results obtained in recovery tests using two water samples.

Sample	$\text{MeHg}^+$ Added/ $\mu\text{g L}^{-1}$	Found in IIP <sup>a</sup> / $\mu\text{g L}^{-1}$	Recovery/%
river water	30.1	$25.5 \pm 0.4$	84.8
	60.4	$53.7 \pm 0.4$	88.8
	114.5	$103 \pm 5$	89.6
tap water	22.4	$21 \pm 2$	93.7
	58.2	$55 \pm 2$	95.2
	113.5	$101 \pm 2$	88.5

<sup>a</sup> Standard deviation for  $n = 3$ .

## 4. Conclusions

The computational simulation based on DFT and experimental validations proved to be a powerful tool to design a  $\text{MeHg}^+$  IIP. The theoretical studies showed that MBI, MBT, AA and EtOH were suitable components for the  $\text{MeHg}$ -IIP synthesis. The performances of the IIP-MBI-AA and IIP-MBT-AA polymers were compared. The effect of pH and adsorption isotherm of imprinted and control polymers revealed differences at  $\text{pH} = 8$  with a superior maximum adsorption capacity for the IIP-MBT-AA material. The adsorption isotherm studies indicated that the Freundlich isotherm fitted well the adsorption of  $\text{MeHg}^+$  on the imprinted and control polymers. Additionally, the kinetic study showed follow the pseudo-second-order kinetic equation well in the adsorption process. The selectivity studies showed that the selectivity coefficient  $k$  values for IIP-MBT-AA were  $\text{MeHg}^+/\text{Hg}^{2+}$  (0.86),  $\text{MeHg}^+/\text{Cd}^{2+}$  (261),  $\text{MeHg}^+/\text{Pb}^{2+}$  (288) and  $\text{MeHg}^+/\text{Zn}^{2+}$  (1510), revealing that mercurial compounds have high affinity for the recognition sites. The IIP-MBT-AA was successfully applied to concentrate  $\text{MeHg}^+$  ions in water samples with satisfactory results.

**Supplementary Materials:** The following are available online at <http://www.mdpi.com/2079-4991/10/12/2541/s1>, Figure S1: FTIR spectra of (A) MBI, IIP-MBI-AA, NIP-MBI-AA, and NIP-AA; and (B) MBT, IIP-MBT-AA, NIP-MBT-AA, and NIP-AA. Figure S2: TGA analysis of IIP-MBI-AA, NIP-MBI-AA, IIP-MBT-AA, NIP-MBT-AA and NIP-AA. Table S1: Comparative selected bond lengths ( $\text{\AA}$ ) for  $\text{MeHg}^+$ -MBI and  $\text{MeHg}^+$ -MBT. Table S2: Predicted binding energy of  $\text{MeHg}^+$ -MBI and  $\text{MeHg}^+$ -MBT with FMs. Table S3: Selectivity parameters for the adsorption of  $\text{Hg}^{2+}$ ,  $\text{Zn}^{2+}$ ,  $\text{Cd}^{2+}$  and  $\text{Pb}^{2+}$  ions with respect to  $\text{MeHg}^+$ .

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